

OBJ: Learn how the charge on the electron was determined MSC: Understanding

6. Which of the following is NOT true regarding cations and anions?
- N^{3-} represents a nitrogen atom that has gained three electrons.
 - An oxygen molecule can lose an electron to form O_2^- .
 - Ca^{2+} represents a calcium atom that has lost two electrons.
 - The formation of a chlorine anion can be written as $\text{Cl} + \text{e}^- \rightarrow \text{Cl}^-$.
 - $\text{N}_2 \rightarrow \text{N}_2^+ + \text{e}^-$ describes a nitrogen molecule forming a +1 cation.

ANS: B DIF: Medium REF: 2.1

OBJ: Write notation for charged species (cations and anions) MSC: Analyzing

7. What is the correct symbol for an α particle?
- ${}^4_2\alpha$
 - ${}^2_2\alpha$
 - ${}^2_4\alpha$
 - ${}^4_0\alpha$
 - ${}^2_0\alpha$

ANS: D DIF: Easy REF: 2.1

OBJ: Define and describe types of radioactivity MSC: Applying

8. What is the correct symbol for a β^- particle?
- ${}^1_0\beta$
 - ${}^{-1}_0\beta$
 - ${}^1_1\beta$
 - ${}^0_1\beta$
 - ${}^0_{-1}\beta$

ANS: E DIF: Easy REF: 2.1

OBJ: Define and describe types of radioactivity MSC: Applying

9. Who discovered neutrons?
- Henri Becquerel
 - Robert Millikan
 - Joseph John Thomson
 - John Dalton
 - James Chadwick

ANS: E DIF: Easy REF: 2.1

OBJ: Describe Rutherford's experiment that discovered the atomic nucleus and the subsequent view of atomic structure MSC: Remembering

10. Which statement regarding Rutherford's nuclear model of the atom is false?
- The diameter of the atom is approximately 10,000 times larger than the nucleus.
 - The nucleus is approximately 100 times smaller than the diameter of the atom.
 - The nucleus is surrounded by a diffuse cloud of electrons.
 - Electrons and protons are not mixed uniformly throughout the atom.
 - The atom is mostly empty space.

ANS: B DIF: Easy REF: 2.1

OBJ: Describe Rutherford's experiment that discovered the atomic nucleus and the subsequent view of atomic structure MSC: Understanding

11. Which statement regarding the Geiger–Marsden experiment is false?
- Beta particles were occasionally deflected by electrons in the gold atoms.
 - Alpha particles were occasionally deflected by small positively charged regions in the gold atoms.
 - The results suggested that the positive charge of an atom is localized in a small region.
 - The results suggested that most of the mass of an atom is contained in a small region.
 - The results suggested that the plum-pudding model of the atom was incorrect.

ANS: A DIF: Easy REF: 2.1

OBJ: Describe Rutherford's experiment that discovered the atomic nucleus and the subsequent view of atomic structure MSC: Understanding

12. If the diameter of a carbon atom is approximately 140 pm, how many carbon atoms lined up side to side would span a pencil lead with a diameter of about 0.7 mm?
- 5×10^1 atoms
 - 5×10^3 atoms
 - 5×10^5 atoms
 - 5×10^6 atoms
 - 5×10^9 atoms

ANS: D DIF: Medium REF: 2.1

OBJ: Describe Rutherford's experiment that discovered the atomic nucleus and the subsequent view of atomic structure MSC: Analyzing

13. The diameter of a carbon atom is approximately 140 pm, whereas the diameter of a pencil lead is approximately 0.7 mm. How many carbon nuclei would be required to span 0.70 mm? The radius of the nucleus is approximately 10,000 times smaller than the radius of an atom.
- 5×10^9 nuclei
 - 5×10^{10} nuclei
 - 5×10^{12} nuclei
 - 5×10^{13} nuclei
 - 5×10^{15} nuclei

ANS: B DIF: Difficult REF: 2.1

OBJ: Describe Rutherford's experiment that discovered the atomic nucleus and the subsequent view of atomic structure MSC: Evaluating

14. A baseball has a diameter of approximately 7.4 cm (2.9 inches), whereas a carbon atom has a diameter of about 140 pm. How many times larger is the baseball than the carbon atom?
- 5.3×10^{12}
 - 5.3×10^{10}
 - 5.3×10^8
 - 5.3×10^6
 - 5.3×10^3

ANS: C DIF: Difficult REF: 2.1

OBJ: Describe Rutherford's experiment that discovered the atomic nucleus and the subsequent view of atomic structure MSC: Evaluating

15. If the nucleus of an atom had a diameter of 1 cm (roughly that of a dime), what would be the approximate diameter of the atom? The radius of the nucleus is approximately 10,000 times smaller than the radius of an atom.
- 1000 km
 - 10 km
 - 1000 m
 - 100 m
 - 1 m

ANS: D DIF: Difficult REF: 2.1

OBJ: Describe Rutherford's experiment that discovered the atomic nucleus and the subsequent view of atomic structure MSC: Evaluating

16. If the nucleus of an atom has a radius of about 5 fm and a mass of about 2×10^{-21} g, what is its approximate density? (Volume of a sphere = $4\pi r^3/3$)
- 4×10^{15} g/cm³
 - 4×10^{12} g/cm³
 - 4×10^9 g/cm³
 - 4×10^6 g/cm³
 - 4×10^3 g/cm³

ANS: A DIF: Difficult REF: 2.1

OBJ: Describe Rutherford's experiment that discovered the atomic nucleus and the subsequent view of atomic structure MSC: Evaluating

17. Which subatomic particles have opposite charges?
- a. protons and neutrons
 - b. protons and electrons
 - c. neutrons and electrons
 - d. all protons
 - e. all neutrons

ANS: B DIF: Easy REF: 2.1

OBJ: Compare the mass and charge of atomic particles: electrons, protons, and neutrons

MSC: Understanding

18. Which subatomic particles have approximately equal masses?
- a. protons and neutrons
 - b. protons and electrons
 - c. neutrons and electrons
 - d. protons, neutrons, and electrons
 - e. none of the above

ANS: A DIF: Easy REF: 2.1

OBJ: Compare the mass and charge of atomic particles: electrons, protons, and neutrons

MSC: Understanding

19. What is the correct symbol for an electron?
- a. ${}^1_1\text{e}$
 - b. ${}^{-1}_1\text{e}$
 - c. ${}^0_1\text{e}$
 - d. ${}^0_0\text{e}$
 - e. ${}^{\bar{0}}_0\text{e}$

ANS: D DIF: Easy REF: 2.1

OBJ: Compare the mass and charge of atomic particles: electrons, protons, and neutrons

MSC: Understanding

20. What is the correct symbol for a proton?
- a. ${}^0_{-1}\text{p}$
 - b. ${}^0_0\text{p}$
 - c. ${}^1_1\text{p}$
 - d. ${}^1_0\text{p}$
 - e. ${}^0_0\text{p}$

ANS: C DIF: Easy REF: 2.1

OBJ: Compare the mass and charge of atomic particles: electrons, protons, and neutrons

MSC: Understanding

21. What is the correct symbol for a neutron?
- a. ${}^1_0\text{n}$
 - b. ${}^1_1\text{n}$
 - c. ${}^0_1\text{n}$
 - d. ${}^1_1\text{n}$
 - e. ${}^{\bar{0}}_0\text{n}$

ANS: A DIF: Easy REF: 2.1

OBJ: Compare the mass and charge of atomic particles: electrons, protons, and neutrons

MSC: Understanding

22. Protons and neutrons are examples of _____
- a. nuclei.
 - b. nuclides.
 - c. nucleons.
 - d. isotopes.
 - e. charged particles.

ANS: C DIF: Easy REF: 2.1

OBJ: Compare the mass and charge of atomic particles: electrons, protons, and neutrons
MSC: Understanding

23. Which of the following statements regarding the discovery of isotopes is false?
- Positively charged ions were deflected by a combination of electric and magnetic fields.
 - Nuclides with equal charges but different masses were deflected to different degrees.
 - The amount of deflection of an ion depended on its charge.
 - An estimate of the relative abundance of the different isotopes of an element could be ascertained.
 - If nuclides had the same mass but different positive charges, the ion with the smallest charge was deflected the most.

ANS: E DIF: Medium REF: 2.2

OBJ: Describe the experiment that discovered isotopes MSC: Analyzing

24. Isotopes have _____
- the same atomic mass.
 - the same total number of protons and neutrons.
 - the same number of neutrons but a different number of protons.
 - the same number of protons but a different number of neutrons.
 - the same number of protons but different numbers of electrons.

ANS: D DIF: Easy REF: 2.2

OBJ: Define isotope, atomic number, mass number MSC: Remembering

25. The ${}^{12}_6\text{C}$ nucleus is an example of _____
- a nuclide.
 - an element.
 - a proton.
 - a neutron.
 - a nucleon.

ANS: A DIF: Easy REF: 2.2

OBJ: Define isotope, atomic number, mass number MSC: Understanding

26. ${}^{12}_6\text{C}$ and ${}^{13}_6\text{C}$ are examples of _____
- ions.
 - neutrons.
 - nucleons.
 - isotopes.
 - charged particles.

ANS: D DIF: Easy REF: 2.2

OBJ: Define isotope, atomic number, mass number MSC: Understanding

27. The atomic model that includes isotopes differs from Dalton's view of the atom. Which of the following statements is false?
- The identity of an atom can be determined solely by its atomic number.
 - The identity of an isotope can be determined solely by its mass number.
 - Atoms of different elements may have the same mass numbers.
 - Atoms of different elements cannot contain the same number of protons.
 - The different isotopes of an element are not always equally abundant.

ANS: B DIF: Medium REF: 2.2

OBJ: Compare Dalton's definition of an element to the definition that realizes the existence of isotopes MSC: Analyzing

28. A ${}^{35}_{17}\text{Cl}$ atom has _____ protons, _____ neutrons, and _____ electrons.
- 17, 18, 17
 - 18, 35, 17

ii	54	131	54
iii	52	128	54

- a. ${}_{56}^{81}\text{Ba}^{2+}$; ${}_{54}^{77}\text{Xe}$; ${}_{52}^{76}\text{Te}^{2+}$
 b. ${}_{56}^{81}\text{Ba}^{2+}$; ${}_{54}^{77}\text{Xe}$; ${}_{52}^{76}\text{Te}^{2+}$
 c. ${}_{56}^{137}\text{Ba}^{2+}$; ${}_{54}^{131}\text{Xe}$; ${}_{52}^{128}\text{Te}^{2+}$
 d. ${}_{56}^{81}\text{Ba}^{2+}$; ${}_{54}^{77}\text{Xe}$; ${}_{52}^{76}\text{Te}^{2+}$
 e. ${}_{56}^{137}\text{Ba}^{2+}$; ${}_{54}^{131}\text{Xe}$; ${}_{52}^{128}\text{Te}^{2+}$

ANS: E DIF: Medium REF: 2.2

OBJ: Interpret and write symbols for nuclides, identify nuclides from mass numbers and atomic numbers, and determine their charges from the number of electrons

MSC: Applying

34. Identify the atom or ion: i) _____; ii) _____; and iii) _____.

<i>Atom</i>	<i>Protons</i>	<i>Neutrons</i>	<i>Electrons</i>
i	17	20	18
ii	18	22	18
iii	19	20	18

- a. ${}_{17}^{35}\text{Cl}^{-}$; ${}_{18}^{40}\text{Ar}$; ${}_{19}^{37}\text{K}^{+}$
 b. ${}_{17}^{37}\text{Cl}^{-}$; ${}_{18}^{40}\text{Ar}$; ${}_{19}^{39}\text{K}^{+}$
 c. ${}_{17}^{37}\text{Cl}^{+}$; ${}_{18}^{40}\text{Ar}$; ${}_{19}^{39}\text{K}^{-}$
 d. ${}_{17}^{35}\text{Cl}^{+}$; ${}_{18}^{40}\text{Ar}$; ${}_{19}^{37}\text{K}^{-}$
 e. ${}_{17}^{35}\text{Cl}^{-}$; ${}_{18}^{40}\text{Ar}$; ${}_{19}^{37}\text{K}^{-}$

ANS: B DIF: Medium REF: 2.2

OBJ: Interpret and write symbols for nuclides, identify nuclides from mass numbers and atomic numbers, and determine their charges from the number of electrons

MSC: Applying

35. What is the nuclide symbol for the atom that has an atomic number equal to the number of electrons in ${}^{35}\text{Cl}^{-}$ and a neutron number equal to the mass number of a sodium atom containing 11 neutrons?

- a. ${}_{18}^{41}\text{Ar}$
 b. ${}_{20}^{42}\text{Ca}$
 c. ${}_{35}^{79}\text{Br}$
 d. ${}_{18}^{40}\text{Ar}$
 e. ${}_{24}^{52}\text{Cr}$

ANS: D DIF: Medium REF: 2.2

OBJ: Interpret and write symbols for nuclides, identify nuclides from mass numbers and atomic numbers, and determine their charges from the number of electrons

MSC: Analyzing

36. What is the nuclide symbol for the ion that has a charge of 2+, 50 neutrons more than its number of protons, and an atomic number equal to the number of electrons in a zirconium atom that has lost 2 electrons?

- a. ${}_{50}^{72}\text{Sn}^{2+}$
 b. ${}_{38}^{88}\text{Sr}^{2+}$
 c. ${}_{40}^{90}\text{Zr}^{2+}$
 d. ${}_{40}^{90}\text{Zr}^{2+}$
 e. ${}_{38}^{90}\text{Sr}^{2+}$

ANS: B DIF: Medium REF: 2.2

OBJ: Interpret and write symbols for nuclides, identify nuclides from mass numbers and atomic numbers, and determine their charges from the number of electrons

MSC: Analyzing

c. -1

ANS: E DIF: Medium REF: 2.3

OBJ: Determine charge from the position of the element in the periodic table

MSC: Analyzing

56. You synthesize a superheavy atom that fits into the periodic table below radium. If it were to form an ion, what ionic charge would you predict?

- a. +2
- b. +1
- c. -1
- d. -2
- e. 0 (unlikely to form an ion)

ANS: A DIF: Difficult REF: 2.3

OBJ: Determine charge from the position of the element in the periodic table

MSC: Creating

57. Which of the following is an alkali metal?

- a. K
- b. Mg
- c. Al
- d. Cu
- e. Ca

ANS: A DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

58. Elements 21 through 30 are known as _____

- a. alkaline earths.
- b. chalcogens.
- c. halides.
- d. transition metals.
- e. rare earths.

ANS: D DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

59. Calcium is an example of _____

- a. an alkali metal.
- b. a transition metal.
- c. an alkaline earth metal.
- d. a halogen.
- e. a chalcogen.

ANS: C DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

60. Elements in group 17 (VIIA) are called _____

- a. alkali metals.
- b. pnictogens.
- c. alkaline earth metal.
- d. halogens.
- e. chalcogens.

ANS: D DIF: Easy REF: 2.4

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

61. Elements in group 18 (VIIIA) are called _____
- a. alkali metals.
 - b. noble gases.
 - c. alkaline earth metals.
 - d. halogens.
 - e. chalcogens.

ANS: B DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

62. Silicon is best described as a _____
- a. metalloid.
 - b. metal.
 - c. transition metal.
 - d. noble gas.
 - e. nonmetal.

ANS: A DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

63. Sodium is best described as a _____
- a. metalloid.
 - b. metal.
 - c. transition metal.
 - d. noble gas.
 - e. nonmetal.

ANS: B DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

64. Cobalt is best described as a _____
- a. metalloid.
 - b. transition metal.
 - c. chalcogen.
 - d. noble gas.
 - e. nonmetal.

ANS: B DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

65. Oxygen is best described as a _____
- a. metalloid.
 - b. metal.
 - c. transition metal.
 - d. noble gas.
 - e. nonmetal.

ANS: E DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

66. What is the name of the metalloid in period four that is in the same family as nitrogen?
- a. bismuth
 - b. antimony
 - c. arsenic
 - d. carbon
 - e. selenium

ANS: C DIF: Medium REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Analyzing

67. What is the name of the halogen in period five?
- a. selenium
 - b. tellurium
 - c. bromine
 - d. iodine
 - e. antimony

ANS: D DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

68. Which halogen is radioactive?
- a. astatine
 - b. polonium
 - c. iodine
 - d. tellurium
 - e. bismuth

ANS: A DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

69. Which characteristic would you expect indium NOT to exhibit?
- a. shiny luster
 - b. electrically insulating
 - c. malleable
 - d. a +3 ionic charge
 - e. solid at room temperature

ANS: B DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

70. You create a superheavy atom with an atomic number of 120. To which category does it belong?
- a. halogens
 - b. actinides
 - c. transition metals
 - d. alkali metals
 - e. alkaline earth metals

ANS: E DIF: Medium REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Analyzing

71. You create a superheavy atom with an atomic number of 120. What is probably true about that element?
- a. It is probably a gas.
 - b. It is probably a metalloid.
 - c. It is probably nonmetallic.
 - d. It is probably metallic.
 - e. It probably forms a stable +1 cation.

ANS: D DIF: Medium REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Analyzing

72. One isotope makes up 97% of all calcium atoms. Which one?

- a. ^{40}Ca
- b. ^{42}Ca
- c. ^{43}Ca
- d. ^{44}Ca
- e. ^{48}Ca

ANS: A DIF: Easy REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Understanding

73. For each of the elements below, there are only two naturally occurring isotopes. Using the atomic masses on the periodic table, identify the pair in which the heavier isotope is the more abundant one.

- a. ^{63}Cu and ^{65}Cu
- b. ^{85}Rb and ^{87}Rb
- c. ^{10}B and ^{11}B
- d. ^{79}Br and ^{81}Br
- e. ^{14}N and ^{15}N

ANS: C DIF: Medium REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Analyzing

74. For each of the elements below, there are only two naturally occurring isotopes. Using the atomic masses on the periodic table, identify the pair in which the lighter isotope is the more abundant one.

- a. ^6Li and ^7Li
- b. ^{79}Br and ^{81}Br
- c. ^{10}B and ^{11}B
- d. ^{121}Sb and ^{123}Sb
- e. ^{50}V and ^{51}V

ANS: B DIF: Medium REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Analyzing

75. Gallium has two naturally occurring isotopes with the following masses and natural abundances. Calculate the average atomic mass of Ga.

^{69}Ga 68.9256 amu 60.108%

^{71}Ga 70.9247 amu 39.892%

- a. 69.925 amu
- b. 70.127 amu
- c. 70.000 amu
- d. 69.824 amu
- e. 69.723 amu

ANS: E DIF: Medium REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Analyzing

76. Rubidium has two naturally occurring isotopes, ^{85}Rb (84.912 amu) and ^{87}Rb (86.909 amu). Rubidium-85 is the more abundant isotope (72.17%). Calculate the average atomic mass of Rb.

- a. 86.91 amu
- b. 85.47 amu
- c. 85.91 amu
- d. 86.35 amu
- e. 86.00 amu

ANS: B DIF: Medium REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Analyzing

77. The average atomic mass of lithium is 6.941 amu. Lithium has two naturally occurring isotopes, ^6Li (7.52%) and ^7Li (92.48%). The mass of ^6Li is 6.0151 amu. What is the isotopic mass of ^7Li ?

- a. 7.016 amu
- b. 0.926 amu
- c. 7.015 amu
- d. 7.000 amu
- e. 6.941 amu

c. 6.001 amu

ANS: A DIF: Difficult REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Evaluating

78. The average atomic mass of silver is 107.868 amu. Silver has two naturally occurring isotopes, ^{107}Ag (106.905 amu, 51.839%) and ^{109}Ag . What is the isotopic mass of ^{109}Ag ?

- a. 109.11 amu
- b. 108.89 amu
- c. 108.52 amu
- d. 108.91 amu
- e. 108.48 amu

ANS: D DIF: Difficult REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Evaluating

79. The average atomic mass of nickel is 58.693 amu. Given the data in the following table, what is the natural abundance of nickel-64?

<i>Isotope</i>	<i>Mass (amu)</i>	<i>Natural Abundance (%)</i>
^{58}Ni	57.935	68.0769
^{60}Ni	59.931	26.2231
^{61}Ni	60.931	1.1399
^{62}Ni	61.928	3.6345
^{64}Ni	63.928	?

- a. 92.56%
- b. 9.256%
- c. 7.440%
- d. 0.9256%
- e. 0.7440%

ANS: D DIF: Medium REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Analyzing

80. Identify the element based on the following values for its three isotopes: 38.9637 amu (93.08%), 39.9640 amu (0.012%), and 40.9618 amu (6.91%).

- a. K
- b. Cl
- c. S
- d. Ar
- e. Ca

ANS: A DIF: Difficult REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Evaluating

81. Identify the element based on the following values for its five isotopes: 179.947 amu (0.12%), 181.948 amu (26.50%), 182.950 amu (14.31%), 183.951 amu (30.64%), and 185.954 amu (28.43%).

- a. Ir
- b. Os
- c. Re
- d. Ta
- e. W

ANS: E DIF: Medium REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Applying

82. Iron has four naturally occurring isotopes, one of which is far more abundant than the others. Iron has an average atomic mass of 55.845 amu. Which isotope is most abundant?
- a. ^{54}Fe , 53.9396 amu
 - b. ^{56}Fe , 55.9349 amu
 - c. ^{57}Fe , 56.9354 amu
 - d. ^{58}Fe , 57.9333 amu
 - e. ^{70}Zn , 69.9253 amu

ANS: B DIF: Difficult REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Evaluating

83. Highly enriched weapons-grade uranium might consist of exactly 90% U-235 (235.044 amu), with the remainder being U-238 (238.051 amu). What is the average atomic mass of this sample of highly enriched uranium? Assume the percentages are exact.
- a. 238.051 amu
 - b. 236.547 amu
 - c. 235.754 amu
 - d. 235.645 amu
 - e. 235.345 amu

ANS: E DIF: Difficult REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Evaluating

84. Weapons-grade plutonium consists of 93% Pu-239 (239.0522 amu), with the remainder being Pu-240 (240.0538 amu). What is the average atomic mass of this sample of weapons-grade plutonium? Assume the percentages are exact.
- a. 239.5530 amu
 - b. 239.1223 amu
 - c. 239.0522 amu
 - d. 239.9478 amu
 - e. 239.9822 amu

ANS: B DIF: Difficult REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Evaluating

85. You synthesize a sample of a superheavy element with an atomic number of 180. In a mass spectrometer, you find two peaks at 5.08113×10^{-25} kg and 5.11434×10^{-25} kg, with the former appearing to be about 2.500 times larger than the latter. What is the approximate average atomic mass of this element in amu? (1 amu = 1.6605×10^{-27} kg)
- a. 475.0 amu
 - b. 306.8 amu
 - c. 306.6 amu
 - d. 430.4 amu
 - e. 1073 amu

ANS: C DIF: Difficult REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Creating

86. What is the molecular mass of phosphorus pentachloride (PCl_5)?
- a. 177.3 amu
 - b. 190.3 amu
 - c. 208.2 amu
 - d. 172.8 amu
 - e. 202.8 amu

ANS: C DIF: Easy REF: 2.4

OBJ: Compute the molecular mass from a formula

MSC: Remembering

87. What is the molecular mass of sulfuric acid (H_2SO_4)?
- a. 49.0 amu
 - d. 98.1 amu

- b. 24.5 amu
c. 101 amu
- e. 97.0 amu

ANS: D DIF: Easy REF: 2.4

OBJ: Compute the molecular mass from a formula

MSC: Remembering

88. What is the formula unit mass of chromium(III) oxide (Cr_2O_3)?
- a. 152.0 amu
b. 136.0 amu
c. 120.0 amu
- d. 104.0 amu
e. 68.0 amu

ANS: A DIF: Easy REF: 2.4

OBJ: Compute the molecular mass from a formula

MSC: Remembering

89. Which has the highest molecular mass?
- a. Br_2O
b. IBr_2
c. CBr_4
- d. Br_2O_8
e. BrF_5

ANS: C DIF: Medium REF: 2.4

OBJ: Compute the molecular mass from a formula

MSC: Analyzing

90. Which has the highest formula unit mass?
- a. CaBr_2
b. NaI_2
c. CdF_2
- d. HgCl_2
e. Ag_2S

ANS: B DIF: Medium REF: 2.4

OBJ: Compute the molecular mass from a formula

MSC: Analyzing

91. Which of the following compounds contains the most nitrogen atoms?
- a. N_2H_4
b. NH_4NO_3
c. NaN_3
- d. $(\text{NH}_4)_2\text{SO}_4$
e. $\text{C}_8\text{H}_{12}\text{N}_4\text{O}_3$

ANS: E DIF: Medium REF: 2.4

OBJ: Compute the molecular mass from a formula

MSC: Analyzing

92. How many CaCl_2 formula units are in 125 amu of calcium chloride?
- a. 0.888 formula units
b. 1.00 formula units
c. 1.13 formula units
- d. 1.25 formula units
e. 3.75 formula units

ANS: C DIF: Medium REF: 2.4

OBJ: Compute the molecular mass from a formula

MSC: Applying

93. How many hydrogen atoms are in 51 amu of pure ammonia (NH_3)?
- a. 3 H atoms
b. 9 H atoms
c. 17 H atoms
- d. 50 H atoms
e. 150 H atoms

ANS: B DIF: Difficult REF: 2.4

OBJ: Compute the molecular mass from a formula

MSC: Evaluating

94. Which contains the most bromine by mass?
- a. Br_2O
d. Br_2O_8

- a. 4.4×10^{17} moles
b. 4.4×10^{-17} moles
c. 1.4×10^{-8} moles
- d. 2.3×10^{-7} moles
e. 7.3×10^{-7} moles

ANS: E DIF: Difficult REF: 2.5

OBJ: Convert between moles and numbers of atoms/molecules MSC: Evaluating

101. Cyanidin chloride ($C_{15}H_{11}O_6Cl$, 322.7 g/mol) contains the cyanidin ion, a pigment found in many berries. Calculate the number of moles of cyanidin chloride equivalent to 7.2 mg.

- a. 2.2×10^1 mol
b. 3.2×10^{-2} mol
c. 2.2×10^{-2} mol
- d. 2.2×10^{-5} mol
e. 7.2×10^{-5} mol

ANS: D DIF: Easy REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Understanding

102. A 1.5 g tablet for pain might contain 0.30 g acetaminophen ($C_8H_9NO_2$, 151.16 g/mol) and 0.044 g codeine ($C_{18}H_{21}NO_3$, 299.36 g/mol). Calculate the number of moles of codeine in the tablet.

- a. 1.5×10^{-4} mol
b. 2.9×10^{-4} mol
c. 1.0×10^{-3} mol
- d. 5.0×10^{-3} mol
e. 1.1×10^{-1} mol

ANS: A DIF: Easy REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Understanding

103. Antirrhinin chloride ($C_{27}H_{11}O_6Cl$, 630.97 g/mol) contains the antirrhinin ion, a pigment found in açai berries. Calculate the number of moles of carbon in 0.75 g of antirrhinin chloride.

- a. 1.2×10^{-3} mol
b. 3.2×10^{-2} mol
c. 4.3×10^{-2} mol
- d. 1.6×10^{-3} mol
e. 4.4×10^{-5} mol

ANS: B DIF: Medium REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Analyzing

104. TNT, or trinitrotoluene, has the chemical formula $C_7H_5N_3O_6$. How many grams of nitrogen are present in 25 grams TNT (227.13 g/mol)?

- a. 0.11 g
b. 0.33 g
c. 1.5 g
- d. 4.6 g
e. 5.4 g

ANS: D DIF: Medium REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Analyzing

105. Calculate the number of americium-241 atoms present in a smoke alarm containing 0.30 μ g of radioactive ^{241}Am (241.06 g/mol).

- a. 7.5×10^{14} atoms
b. 1.8×10^{17} atoms
c. 2.5×10^{21} atoms
- d. 6.0×10^{23} atoms
e. 1.4×10^5 atoms

ANS: A DIF: Medium REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Applying

106. Gold prices in early 2013 were approximately \$51 per gram. Given that the mass of a gold atom is approximately 3.27×10^{-22} grams, how many gold atoms could you buy with a quarter (\$0.25)?
- 7.6×10^{20} atoms
 - 1.5×10^{19} atoms
 - 6.5×10^{19} atoms
 - 1.2×10^{22} atoms
 - 1.1×10^{23} atoms

ANS: B DIF: Difficult REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Evaluating

107. In early 2013, platinum (195.078 g/mol) was selling for about \$1575 per troy ounce (one troy ounce is equal to 31.10 g). How many platinum atoms could you buy for \$20.00?
- 1.219×10^{21} atoms
 - 7.817×10^{21} atoms
 - 7.587×10^{21} atoms
 - 2.360×10^{23} atoms
 - 2.431×10^{23} atoms

ANS: A DIF: Difficult REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Evaluating

108. In the presence of cyanuric acid ($C_3H_3N_3O_3$), melamine ($C_3H_6N_6$) can form crystals that potentially cause renal failure if ingested. Which contains the most nitrogen?
- 1.22×10^{24} $C_3H_6N_6$ molecules
 - 2.78 mol $C_3H_6N_6$
 - 6.54×10^{24} $C_3H_3N_3O_3$ molecules
 - 10.3 mol $C_3H_3N_3O_3$
 - 2.56 mol $C_3H_6N_6$ + 2.56 mol $C_3H_3N_3O_3$

ANS: C DIF: Difficult REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Evaluating

109. TNT, or trinitrotoluene, has the chemical formula $C_7H_5N_3O_6$ (227.13 g/mol). Which of the following amounts of TNT contains the most nitrogen?
- 1.22×10^{24} molecules
 - 278 g
 - 1.33 mol
 - the number of moles of TNT containing 9 moles of O
 - the number of moles of TNT containing 8 moles of C

ANS: A DIF: Difficult REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Evaluating

110. TNT ($C_7H_5N_3O_6$, 227.13 g/mol) and RDX ($C_3H_6N_6O_6$, 222.12 g/mol) are explosive materials. Which of the following contains the most nitrogen?
- 3.45×10^{23} molecules RDX
 - 145 g TNT
 - 95.4 g RDX
 - 4.25×10^{23} molecules TNT
 - 0.875 mol RDX

ANS: E DIF: Difficult REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Evaluating

111. How many grams of P_4 (123.88 g/mol) would contain the same number of atoms as 154 g S_8 (256.48 g/mol)?

MSC: Understanding

117. According to the Big Bang theory, which statement about the origin of the elements is NOT correct?
- Initially, energy was transformed into electrons and other elementary particles.
 - As the universe cooled, neutrons and protons were formed.
 - Collisions of neutrons and protons produced deuterons, which then led to the formation of alpha particles.
 - Nuclear fusion reactions in the interior of stars formed elements up to ^{56}Fe .
 - All nuclear reactions forming the elements required an input of energy.

ANS: E DIF: Medium REF: 2.6

OBJ: Describe the nucleosynthesis of the elements up to uranium after the Big Bang

MSC: Analyzing

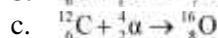
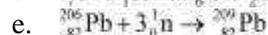
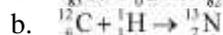
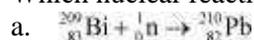
118. Which statement is NOT correct regarding primordial nucleosynthesis?
- Colliding pairs of electrons annihilated each other to form two gamma rays.
 - Deuterons fused together, forming alpha particles.
 - More stable nuclides were formed from less stable nuclides.
 - Gamma rays were produced.
 - Neutrons and protons fused together, forming deuterons.

ANS: A DIF: Medium REF: 2.6

OBJ: Describe the nucleosynthesis of the elements up to uranium after the Big Bang

MSC: Analyzing

119. Which nuclear reaction is NOT correctly written?

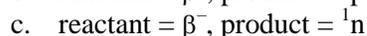
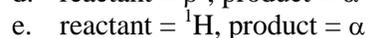
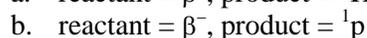
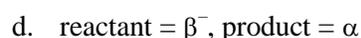
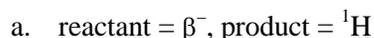
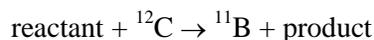


ANS: A DIF: Medium REF: 2.7

OBJ: Write and balance nuclear reactions

MSC: Analyzing

120. Which reactant and product would balance the following nuclear reaction equation?



ANS: C DIF: Difficult REF: 2.7

OBJ: Write and balance nuclear reactions

MSC: Evaluating

121. The peak in nuclear binding energy/nucleon occurs at an isotope of _____

a. helium.

d. carbon.

b. iron.

e. lead.

c. uranium.

ANS: B DIF: Easy REF: 2.7

OBJ: Use the mass defect to calculate the binding energy of a nucleus and the binding energy per nucleon

MSC: Remembering

122. Calculate the nuclear binding energy of the ^{56}Fe nucleus given the following data:

^{56}Fe nuclear mass 55.920679 amu 9.285846×10^{-26} kg

Proton mass 1.00727646 amu 1.672622×10^{-27} kg

Neutron mass 1.00866492 amu 1.674927×10^{-27} kg

Speed of light 2.998×10^8 m/s

- a. 8.346×10^{-9} J
b. 4.417×10^{-9} J
c. 4.370×10^{-9} J
d. 7.804×10^{-11} J
e. 7.887×10^{-11} J

ANS: E DIF: Difficult REF: 2.6

OBJ: Use the mass defect to calculate the binding energy of a nucleus and the binding energy per nucleon
MSC: Evaluating

123. Which of the following statements regarding the mass defect and nuclear binding energy is false?
- The mass of the nucleus is slightly less than the combined mass of its separate constituent nucleons.
 - Separated protons and neutrons are more stable than when they are in the nucleus.
 - Mass is converted to energy when separated nucleons combine to form a nucleus.
 - The binding energy reflects the amount of energy that would be required to break up the nucleus.
 - $E = mc^2$ allows binding energies to be calculated.

ANS: B DIF: Medium REF: 2.7

OBJ: Use the mass defect to calculate the binding energy of a nucleus and the binding energy per nucleon
MSC: Analyzing

124. Which statement regarding the strong nuclear force is false?
- The strong nuclear force prevents radioactive decay from occurring.
 - The strong nuclear force is about 100 times stronger than the repulsive force between protons.
 - Nuclear stability depends on the competition between the strong nuclear force and electrostatic repulsions.
 - The strong nuclear force acts only over very short distances.
 - The strong nuclear force binds nucleons together and stabilizes the nucleus.

ANS: A DIF: Easy REF: 2.7

OBJ: Describe the strong nuclear force MSC: Understanding

125. A supernova event is the explosion caused by the collapse of a dying star that has run out of its nuclear fuel. These stars and events are responsible for _____
- the production of elements heavier than iron-56.
 - nuclear fission of heavy elements.
 - the distribution of heavy elements throughout the universe.
 - both a and c.
 - both b and c.

ANS: D DIF: Medium REF: 2.7

OBJ: Describe the role supernovas have in nucleosynthesis MSC: Applying

126. Elements higher than uranium in the periodic table must be synthesized. Identify the nuclear synthesis reaction that is NOT correctly written.

- a. ${}_{94}^{244}\text{Pu} + {}_{20}^{48}\text{Ca} \rightarrow {}_{114}^{289}\text{Fl} + 3{}_0^1\text{n}$
b. ${}_{98}^{249}\text{Cf} + {}_7^{14}\text{N} \rightarrow {}_{105}^{260}\text{Db} + 4{}_0^1\text{n}$
c. ${}_{83}^{209}\text{Bi} + {}_{24}^{54}\text{Cr} \rightarrow {}_{107}^{262}\text{Bh} + {}_0^1\text{n}$
d. ${}_{92}^{238}\text{U} + {}_{20}^{48}\text{Ca} \rightarrow {}_{112}^{282}\text{Cn} + 4{}_0^1\text{n}$
e. ${}_{96}^{248}\text{Cm} + {}_{20}^{48}\text{Ca} \rightarrow {}_{116}^{293}\text{Cn} + 3{}_0^1\text{n}$

ANS: B DIF: Medium REF: 2.7

OBJ: Describe how elements heavier than uranium are synthesized
MSC: Analyzing

5. Write the complete atomic symbol with both a superscript and a subscript for the atom or ion that contains 11 protons, 10 electrons, and 12 neutrons.

ANS:



DIF: Easy REF: 2.2

OBJ: Interpret and write symbols for nuclides, identify nuclides from mass numbers and atomic numbers, and determine their charges from the number of electrons

MSC: Understanding

6. Write the complete atomic symbol with both a superscript and a subscript for the atom or ion that contains the same number of protons as the number of neutrons in ${}^{56}\text{Fe}$, has a +2 charge, and has a mass number that equals the atomic number of terbium, Tb.

ANS:



DIF: Medium REF: 2.2

OBJ: Interpret and write symbols for nuclides, identify nuclides from mass numbers and atomic numbers, and determine their charges from the number of electrons

MSC: Analyzing

7. Write the complete atomic symbol with both a superscript and a subscript for the atom or ion that contains the same number of electrons as argon, has a -2 charge, and contains equal numbers of protons and neutrons.

ANS:



DIF: Medium REF: 2.2

OBJ: Interpret and write symbols for nuclides, identify nuclides from mass numbers and atomic numbers, and determine their charges from the number of electrons

MSC: Analyzing

8. A cation has a _____ charge because it has _____ electrons.

ANS:

positive/lost

DIF: Easy REF: 2.3

OBJ: Determine charge from the position of the element in the periodic table

MSC: Understanding

9. An anion has a _____ charge because it has _____ electrons.

ANS:

negative/gained

DIF: Easy REF: 2.3

OBJ: Determine charge from the position of the element in the periodic table

MSC: Understanding

10. What is the charge on an alkali metal atom when it is in an ionic compound?

ANS:

+1

DIF: Easy REF: 2.3

OBJ: Determine charge from the position of the element in the periodic table

MSC: Remembering

11. What is the charge on the phosphorus atom when it forms an ionic compound with magnesium?

ANS:

-3

DIF: Easy REF: 2.3

OBJ: Determine charge from the position of the element in the periodic table

MSC: Remembering

12. What is the charge on the copper ion in CuCl_2 ?

ANS:

+2

DIF: Medium REF: 2.3

OBJ: Determine charge from the position of the element in the periodic table

MSC: Analyzing

13. How many nitrogen atoms would be required to form an ionic compound with barium?

ANS:

2

DIF: Medium REF: 2.3

OBJ: Determine charge from the position of the element in the periodic table

MSC: Analyzing

14. How many oxygen atoms would be required to form an ionic compound with aluminum?

ANS:

3

DIF: Medium REF: 2.3

OBJ: Determine charge from the position of the element in the periodic table

MSC: Analyzing

15. Give an example of an alkali metal.

ANS:

lithium, sodium, potassium, rubidium, cesium, francium; answers will vary.

DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

16. Give an example of an alkaline earth metal.

ANS:

beryllium, magnesium, calcium, strontium, barium, radium; answers will vary.

DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

17. Give an example of a halogen.

ANS:

fluorine, chlorine, bromine, iodine, astatine; answers will vary.

DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

18. Give an example of a period 4 transition metal.

ANS:

scandium, titanium, vanadium, chromium, manganese, iron, cobalt, nickel, copper, zinc; answers will vary.

DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

19. Give an example of a nonmetal.

ANS:

carbon, nitrogen, oxygen, fluorine, sulfur, etc.; answers will vary.

DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

20. Give an example of a metalloid (also known as a semimetal).

ANS:

boron, silicon, germanium, arsenic, antimony, tellurium; answers will vary.

DIF: Easy REF: 2.3

OBJ: Identify on the periodic table: groups, periods, metals, metalloids, nonmetals, representative elements, transition metals, alkali metals, alkaline earth metals, halogens, noble gases

MSC: Remembering

21. What is the average atomic mass of a sample of highly enriched uranium uranium that contains exactly 20% uranium-235 (235.04 amu) and 80% uranium-238 (238.05 amu)?

ANS:
237.45 amu

$$(0.20)(235.04) + (0.80)(238.050) = 237.45 \text{ amu}$$

DIF: Difficult REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Evaluating

22. Boron, which has an average atomic mass of 10.81 amu, has two stable isotopes: boron-10 and boron-11. Boron-10 has an atomic mass of 10.0129 amu and a natural abundance of 19.78%. What is the atomic mass of boron-11?

ANS:
11.01 amu

$$(0.1978)(10.0129) + (0.8022)(x) = 10.81 \text{ amu}$$

DIF: Difficult REF: 2.4

OBJ: Compute the average atomic masses using natural abundances of isotopes for an element

MSC: Evaluating

23. What is the mass in amu of one molecule of glucose, $C_6H_{12}O_6$?

ANS:
180.16 amu

DIF: Easy

REF: 2.5

OBJ: Compute the molecular mass from a formula

MSC: Understanding

24. How many atoms are there in 2.5 moles of water?

ANS:
 4.5×10^{24}
 $2.5 \times 3 \times 6.02 \times 10^{23} = 4.5 \times 10^{24}$

DIF: Medium REF: 2.5

OBJ: Convert between moles and numbers of atoms/molecules MSC: Applying

25. How many hydrogen atoms are there in 473 g of water (roughly 16 fluid ounces)?

ANS:
 3.16×10^{25}
 $473 / 18.02 \times 6.02 \times 10^{23} \times 2 = 3.16 \times 10^{25}$

DIF: Medium REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Applying

26. Using a scanning tunneling microscope, Don Eigler at IBM arranged 99 iron atoms on a copper surface to form the Kanji characters for "atom." What is the total mass of iron present in grams and in atomic mass units? Assume that the Fe atoms are "average" in terms of their mass.

ANS:

9.182×10^{-21} g; 5529 amu

DIF: Medium REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Analyzing

27. Fill in the following table.

Substance	Mass (g)	Moles (mol)	#molecules	#atoms
NO	64.0			
NO ₂		0.786		
N ₂ O			7.52×10^{21}	
N ₂ O ₄				1.48×10^{26}

ANS:

Substance	Mass (g)	Moles (mol)	#molecules	#atoms
NO	64.0	2.13	1.28×10^{24}	2.56×10^{24}
NO ₂	36.2	0.786	4.73×10^{23}	1.42×10^{24}
N ₂ O	0.550	0.0125	7.52×10^{21}	2.26×10^{22}
N ₂ O ₄	3770	41.0	2.47×10^{25}	1.48×10^{26}

DIF: Difficult REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Evaluating

28. Hydrogen sulfide (H₂S) is a highly toxic gas that smells like rotten eggs. Suppose the odor detection limit is approximately 4.7×10^{-7} g H₂S per one gram of air. At this level, how many moles of H₂S are present in 1.0 L air? How many H₂S molecules? Assume the density of air is 0.0013 g/mL.

ANS:

1.8×10^{-8} moles; 1.1×10^{16} molecules.

DIF: Difficult REF: 2.5

OBJ: Use molar mass to convert between mass and moles of a substance

MSC: Evaluating

29. Calculate the formula unit mass of sodium phosphate (Na₃PO₄) in which all of the phosphorus is ³²P, a radioactive isotope of phosphorus used in medical applications. A ³²P atom has an atomic mass of 31.97 amu.

ANS:

164.94 g/mol

$$3(22.99) + 1(31.974) + 4(16.00) = 164.94$$

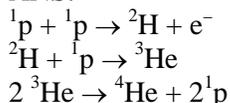
DIF: Medium REF: 2.5

OBJ: Determine the molar mass/formula mass of a substance using the periodic table

MSC: Analyzing

30. Write nuclear reaction equations to show how helium-4 nuclides are produced from protons in our Sun using the following information: step (1) 2 protons react to form hydrogen-2 and a high-energy electron; step (2) hydrogen-2 reacts with a proton to form helium-3; and step (3) two helium-3 combine to helium-4 and two protons.

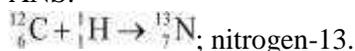
ANS:



DIF: Medium REF: 2.6 OBJ: Write and balance nuclear reactions
 MSC: Analyzing

31. The carbon-nitrogen-oxygen cycle in stars is one process by which hydrogen is converted to helium. Write the complete nuclear reaction for the reaction of a carbon-12 nucleus with a hydrogen nucleus to form nitrogen. What isotope of nitrogen is produced?

ANS:



DIF: Medium REF: 2.6 OBJ: Write and balance nuclear reactions
 MSC: Applying

32. Suppose the reaction ${}^{13}_6\text{C} + {}^1_1\text{H} \rightarrow {}^{14}_7\text{N}$ produces 1.21×10^{-12} J of energy ($1.21 \times 10^{-12} \text{ kg} \cdot \text{m}^2/\text{s}^2$). Calculate the change in mass that occurs during the reaction in amu. $E = mc^2$, where $c = 2.998 \times 10^8$ m/s;
 1 kg = 6.0221415×10^{26} amu.

ANS:

0.00811 amu

$$1.21 \times 10^{-12} \text{ kg} \cdot \text{m}^2/\text{s}^2 / (2.998 \times 10^8 \text{ m/s})^2 \times 6.0221415 \times 10^{26} = 0.00811 \text{ amu}$$

DIF: Difficult REF: 2.6
 OBJ: Use the mass defect to calculate the binding energy of a nucleus and the binding energy per nucleon
 MSC: Evaluating

33. Calculate the binding energy of a helium-4 nucleus in J/mol ${}^4_2\text{He}$. 1 amu = 1.6605×10^{-27} kg

helium-4 nucleus	4.00153 amu
proton mass	1.00728 amu
neutron mass	1.00866 amu
speed of light	2.998×10^8 m/s

ANS:

2.728×10^{12} J/mol

DIF: Difficult REF: 2.6
 OBJ: Use the mass defect to calculate the binding energy of a nucleus and the binding energy per nucleon
 MSC: Evaluating

34. Iron-56 has one of the highest binding energies of all nuclides. Calculate its nuclear binding energy in kJ per mol nucleon. 1 amu is equivalent to 1.492×10^{-10} J.

mass of iron-56 nuclide	55.934994 amu (includes electrons)
proton mass	1.00728 amu
neutron mass	1.00866 amu
electron mass	5.4858×10^{-4} amu
speed of light	2.998×10^8 m/s

ANS:

$$8.477 \times 10^8 \text{ kJ/mol nucleon}$$

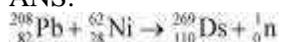
$$56.44908 - 55.92073 = 0.52835 \text{ amu} \times 1.492 \times 10^{-10} \text{ J} \times 6.022 \times 10^{23} / 1000 / 56 = 8.477 \times 10^8 \text{ kJ/mol nucleon}$$

DIF: Difficult REF: 2.6

OBJ: Use the mass defect to calculate the binding energy of a nucleus and the binding energy per nucleon MSC: Evaluating

35. Darmstadtium was first created in 1994 when ^{208}Pb was bombarded with ^{62}Ni to produce ^{269}Ds and one neutron. Write the complete nuclear equation.

ANS:



DIF: Easy REF: 2.7

OBJ: Describe how elements heavier than uranium are synthesized

MSC: Understanding